MULTIPLE CHOICE QUESTION:-

Q.1. Which of the following is a displacement reaction?

(a) $MgCO_3$-- $MgO+CO_2$

(b) $2Na+2H_2O$-- $2NaOH+H_2$

(c) $2H_2+O_2$--- $2H_2O$

(d) $2Pb(NO_3)_2$-- $2PbO+4NO_2+O_2$

(Reason:- Here, sodium displaces to form sodium hydroxide)

Q.2. Magnesium ribbon is rubbed before burning because it has a coating of:-

(a) Basic magnesium carbonate

(b) Basic magnesium oxide

(c) Basic magnesium sulphide

(d) Basic magnesium chloride

Q.3. Which of the following are exothermic processes?

(a) Reaction of water with quick lime

(b) Dilution of an acid

(c) Evaporation of water

(d) Sublimation of camphor

(Reason: - In both the cases heat energy is evolved)

Q.4. Oxidation is a process which involves:

(a) addition of hydrogen

(b) addition of oxygen

(c) removal of oxygen

(d) removal of hydrogen

Q.5. The process of reduction involves-

(a) addition of oxygen

(b) addition of hydrogen

(c) removal of oxygen

(d) removal of hydrogen

Q.6. Give the ratio in which hydrogen and oxygen are present in water by volume?

(a) 1:2

(b) 1:1

(c)2:1

(d) 1:8

Q.7. MnO_2+4HCl--- $MnCl_2+2H_2O+Cl_2$

Identify the substance oxidized in the above equation

(a) $MnCl_2$
(b) HCl
(c) H_2O
(d) MnO_2

(Reason: In this reaction HCl is oxidized to Cl_2, whereas MnO_2 is reduced to $MnCl_2$)

Q.8. When Ag is exposed to air it gets a black coating of

(a)$AgNO_3$

(b)Ag_2S

(c) Ag_2O
(d) Ag_2CO_3
Q.9. Which of the following is an endothermic process?
(a) Dilution of sulphuric acid
(b) Sublimation of dry ice
(c)Condensation of water vapors
(d) Respiration in human beings
Q.10. What type of chemical reaction take place when electricity is passed through water?
(a) Displacement
(b) Combination
(c)Decomposition
(d) Double displacement
Q.11. A substance added to food containing fats and oils is called
(a) Oxidant
(b) Rancid
(c) Coolant
(d) Antioxidant

Q.12. The condition produced by aerial oxidation of fats and oils in foods marked by unpleasant smell and taste is called

(a) Antioxidation

(b) Reduction

(c) Rancidity

(d) Corrosion

Q.13. Electrolysis of water is a decomposition reaction. The mole ratio of hydrogen and oxygen gases liberated during electrolysis of water is

(a) 1:1

(b) 2:1

(c) 4:1

(d) 1:2

Q.14. Which of the following gases can be used for storage?

(a) CO_2 or O_2

(b) Nitrogen or Helium

(c) Helium or Nitrogen

(d) CO_2 or Helium

Q.15. In the reaction $Hg_2Cl_2+Cl_2$- $2HgCl_2$. The reducing agent is

(a) Hg_2Cl_2

(b) Cl_2

(c) $HgCl_2$

(d) Both Cl_2 and $HgCl_2$

Q.16. The brown gas evolved on heating of copper nitrate is

(a) O_2

(b) NO_2

(c) N_2

(d) NO

Q.17. $2AgI\ (s)$- $2Ag(s)+I_2(g)$

The color of Iodine is-

(a) Green

(b) Purple

(c) Brown

(d) Pink

(Iodine is purple in color)

Q.18. Which of the following does not corrode when exposed to the atmosphere?

(a) Iron

(b) Copper

(c) Gold

(d)Silver

Q.19. Black and white photography uses

(a) Decomposition of silver chloride

(b) Decomposition of silver bromide

(c) both

(d) none of these

Q.20.Combination of phosphorous and oxygen is an example of

(a) Oxidation

(b) Reduction

(c) Rancidity

(d) None of these